

A) extraction of hydrogen chloride from products of chlorination of organic compounds (chlorination, dehydrochlorination, pyrolysis, etc.);

- B) synthesis from metal chlorides;
- C) isolation from hydrochloric acid;
- D) thermal decomposition of chlorides;
- E) synthesis by electrolysis of table salt.

7. The electrolysis of the NaCl solution occurs with the formation of:

- A) sodium;
- B) oxygen;
- C) water;
- D) hydrogen;
- E) sodium oxide.

8. The electrolysis of the NaCl solution occurs with the formation of:

- A) sodium;
- B) oxygen;
- C) chlorine;
- D) sodium hypochlorite;
- E) sodium oxide.

9. The electrolysis of the NaCl solution occurs with the formation of:

- A) sodium;
- B) oxygen;
- C) water;
- D) sodium hypochlorite.
- E) sodium hydroxide.

10. The current output is the ratio of the mass of the substance actually obtained by electrolysis to the mass of the substance:

- A) which should be obtained according to the equation of chemical reactions;
- B) electrolyte;
- C) which is consumed in the process;
- D) which should be obtained according to Faraday's law;
- E) empirically calculated.

11. The current output is the ratio of the mass of the substance actually obtained by electrolysis to the mass of the substance:

- A) electrolyte;
- B) which is consumed in the process;
- C) empirically calculated;
- D) according to the estimated quantity;
- E) after electrolysis.

12. The current output is the ratio of the mass of the substance actually obtained by electrolysis to the mass of the substance:

- A) which should be obtained according to the chemical reaction equation;
- B) electrolyte;
- C) empirically calculated;
- D) theoretically calculated;
- E) after electrolysis.